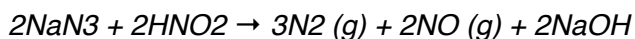
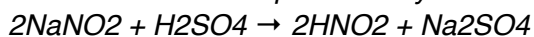


For larger scale work with azide, a good and easy way to quench large scale N_3^- .

" N_3H or N_3^- can be quenched by nitrous acid:



This operation should only be carried out in a fume hood using a vessel with a gas outlet:

- 1 In a three-neck flask with a stirrer, dilute the azide with water so that it does not exceed 5%*
- 2 With stirring add a solution of sodium nitrite containing 1.5 g of sodium nitrite per gram of excess azide*
- 3 Add 2-3M H_2SO_4 dropwise (on small scale, by pipette; on large scale, using a dropping funnel) until gas evolution has ceased and the solution is acidic by test paper.*

Iodide-starch paper can be used to test for an excess of nitrite (goes blue) which indicates that the quench is complete. Dispose of the solution as you would other aqueous waste."

$\text{NO}(\text{g})$ is also a little nasty but very preferable to $\text{HN}_3(\text{g})$! It will just bubble out into the fume hood and go away.

Also of note: don't extract water-azide solutions with DCM or CHCl_3 - these can actually react with azide to form explosive compounds.